1. SHNAYDER, YE.YE.

2. USSR (600)

4. Sugar Machinery

7. "Work of a preheater with small diameter pipes"(B. Zimmermann, "Listy Cukrovarnick vol. 68, no. 6, 1952, p. 140) Sakh.prom. 26 no. 12, 1952

9. Monthly List of Russian Accessions, Library of Congress, February 1953. Unclassified.

SHNAYDER, Ye.Ye.

One-column diffusion past and present [Listy Cukrovarnicke 70 no.7 '54]. Reviewed by E.E.Shnaider. Sakh.prom. 29 no.3:40 '55. (Diffusers) (Sugar industry--Equipment and supplies) (MLRA 8:8)

SHNAYDER, Ye.Ye.

The Czechoslovak journal "Listy Cukrovarnicke" in 1954. Sakh. prom. 29 no.3:42-44 '55. (MLRA 8:7)
(Czechoslovakia--Sugar industry--Periodicals)

SHNAYDER, Ye.Ye.

"Ionites, their properties and use [in Czhech]. J.Smid. Reviewed by E.E.Shnaider. Sakh.prom. 29 no.4:47 '55. (MLRA 8:9)
(Ion exchange) (Smid,J.)

SHNAYDER, Ye.Ye.

Methods for increasing the sugar yield from massecuite and lowering the number of boilings of products. ("Listy cukrovarnicke", no.6, 1955) Abstracted by E.E.Shnaider. Sakh.prom.30 no.5:67 My '56. (MIRA 9:9)
(Sugar industry)

SHNAYDER, Ye.Ye.

New developments in the production of liquid sugar in the U.S.A. ("This is liquid Sugar". 1955; "Zeitschrift für die Zuckerindustrie", no.12, 1955). Abstracted by E.E. Shnaider. Sakh.prom. 30 no.5:73-74 My '56. (Syrups) (MLRA 9:9)

SHNAYDER, Ye.Ye.

Extraction of sugar out of sugar-beets (Listy cukrovarnicke" no.7 1955). Abstracted by E.E.Shnaider. Sakh. prom. 30 no.5:74 My '56. (MLRA 9:9)
(Sugar industry)

SHNAYDER, Ye.ye.

The loss of sugar in the unused sugar beet tailings and small pieces. ("Listy cukrovarnicke", no.1 1956). Abstracted by E.E.Shnaider. Sakh. prom. 30 no.5:74 My '56. (MLRA 9:9)
(Sugar industry--By-products)

SHNAYDER, Ye.Ye.

Observations on the operation of Olier diffusion apparatus ("Zeitschrift für die Zuckerindustrie" no.11 1955). Sakh. prom. 30 no.8:72 Ag. '56. (MLRA 9:11)
(Diffusers)

SHNAYDER, Ye.Ye.

Sugar campaign of 1955-56 in Czechoslovakia ("Listy Cukrovarnicke" no.4 1956). Sakh.prom. 30 no.8:75 Ag. '56. (MLRA 9:11)
(Czechoslovakia--Sugar industry)

SHNAYDER, Ye.Ye.

Characteristics of sugar-cane production (From "Listy cukrovarnicke," no.2 1956) [Reviewed by E.E. Shnaider]. Sakh.prom. 30 no.9:78-79 S '56. (MLRA 10:3)

(Sugar cane)

SHNAYDER, Ye.Ye.

Microbiological investigation of sugar (from "Listy Cucrovarnicke," no.3 1957). Reviewed by E.E. Shnaider. Sakh. prom. 31 no.11:72 N '57. (MIRA 11:1)

(Sugar--Bacteriology)

SHNAYDER, Ye.Ye.

Determining the decolorizing capacity of bone char (from "Listy Cucrovarnicke," no.7 1957). Reviewed by E.E. Shnaider. Sakh. prom. 31 no.11:72 N '57. (MIRA 11:1)

(Animal charcoal)

SHNAYDER, YE.YE.

SHNAYDER, Ye.Ye.

Effect of the magnetic field on scale formation (from "Listy Cucrovarnicke," no.12 1956). Reviewed by E.E. Shnaider. Sakh. prom. 31 no.11:72-73 N '57. (MIRA 11:1)
(Evaporating appliances)

SHPUNTOVA, M.Ye.; SHNAYDER, Ye.Ye.; CHEPIGO, S.V.

Combinated hydrolysis of vegetable matter by concentrated sulfuric acid. Uzb. khim. zhur. no.381-92 '58. (MIRA 11:9)

1.Vsesoyuznyy nauchno-issledovatel'skiy institut sul'fitnospirtovoy i gidroliznoy promyshlennosti.
(Lignin) (Hydrolysis) (Sulfuric acid)

SHNAYDER, Ye.Ye.

Determining moisture loss in white granulated sugar(from "Listy cukrovarnicke," no. 1, 1958). Sakh. prom. 32 no. 6:74 Je '58. (MIRA 11:7)

(Sugar--Storage)

SHNAYDER, Ye.Ye.

Polyelectrolytes as coagulants (from "Listy cucrovarnicke", nos. 4 and 5 1958). Sakh. prom. 32 no.8:69-70 Ag '58. (MIRA 11:9)
(Electrolytes) (Sugar manufacture)

SHNAYDER, Ye.Ye.

Ammonia as fertilizer for sugar beets (from "Listy Cukrovarnicke" no.1, 1958) Sakh.prom. 32 no. 9:72 S '58. (MIRA 11:11)
(Czechoslovakia--Ammonia) (Czechoslovakia--Sugar beets)

SHNAYDER, Ye.Ye.

Boiling out of the evaporator (from "Listy Cukrovarnicke, "No.5, 1958). Sakh.prom. 32 no.12:56 D '58. (MIRA 11:12)
(Czechoslovakia--Sugar industry--Equipment and supplies)
(Corrosion and anticorrosives)

SHNAYDER, Ye.Ye.

Almanac of the sugar beet planter. Sakh. prom. 32 no.12:63 D '58.
(MIRA 11:12)

(Czechoslovakia--Sugar growing)

SHNAYDER, Ye.Ye.

Cold desugarization of green sirup (from "Listy Cukrovarnicke, " No.6, 1958). Sakh. prom. 33 no.2:71-72 F '59. (MIRA 12:3)
(Czechoslovakia--Sugar manufacture)

SHNAYDER, Ye.Ye.

Filtration of saturation juice (from "Chem. zvésti," No.3, 1958). Sakh. prom. 33 no.4:70 Ap '59. (MIRA 12:6)
(Sugar manufacture)

SHNAYDER, Ye.Ye.

Trip to Austria (from "Listy cukrovarnické," No.6, 1958). Sakh. prom. 33 no.4:71-72 Ap '59. (MIRA 12:6)
(Austria--Sugar beets)

SHNAYDER, Ye.Ye.

Works of the Rumanian Scientific Research Institute of the Food Industry, 1958, vol. 2. Sakh. prom. 33 no.4:77 Ap '59.
(MIRA 12:6)

(Sugar manufacture)

SHNAYDER, Ye.Ye.; SHPUNTOVA, M.Ye.; CHEPIGO, S.V.

Combined method of corncob hydrolysis with concentrated sulfuric acid. Gidroliz. i lesokhim.prom. 13 no.7:1-4 '60. (MIRA 13:10)

1. Nauchno-issledovatel'skiy institut gidroliznoy i sul'fitno-spirtovoy promyshlennosti.
(Corncobs) (Hydrolysis)

SHNAYDER, Ye.Ye.

Investigating the particles of a deposit obtained in the first carbonation by different methods of juice purification. Sakh.prom. 34 no.9:72-73 S '60. (MIRA 13:9)
(Sugar manufacture)

ODINTSOV, P.N.; KALNIN'SH, A.I. [Kalnins, A.]; KAL'NINA, V.K.; CHEPIGO, S.V.; SHNAYDER, Ye.Ye.; SHPUNTOVA, M.Ye.

Hydrolysis of plant materials by concentrated sulfuric acid. Gidroliz. i lesokhim.prom. 14 no.3:1-4 '61. (MIRA 14:4)

1. Institut lesokhozyaystvennykh problem i khimii drevesiny Akademii nauk Latviyskoy SSR (for Odintsov, Kalnin'sh, Kal'nina). 2. Nauchno-issledovatel'skiy institut gidroliznoy i sul'fitno spirtovoy promyshlennosti (for Chepigo, Shnayder and Shpuntova).
(Hydrolysis) (Wood--Chemistry)

SHNAYDER, Ye. Ye.

Testing of a pneumatic dryer for potato starch (from "Prumysl Potravin," 12, no.7, 1961). Sakh. prom. 36 no.10:71 O '62. (MIRA 15:10)

(Czechoslovakia--Drying apparatus--Testing)

SHPUNTOVA, M.Ye.; SHNAYDER, Ye.Ye.; CHEPUGO, S.V.; LAZAREVA, L.V.; MASLOVA, L.G.; ROSHCHINA, V.I.; Prinimali uchastiye: PAVLENKO, V.M., starshiy laborant; GERASIMOVA, L.I., starshiy laborant

Pentose hydrolysis of cottonseed hulls and corncobs with hexose hydrolyzates. Sbor.trud. NIIGS 11:7-15 '63. (MIRA 16:12)

NAYDENOV, A.K.; SHNAYDER, Ye.Ye.; SHPUNTOVA, M.Ye.

Dryer for cellolignin obtained from corncobs. Gidroliz. i lesokhim. prom. 16 no.6:7-10 '63. (MIRA 16:10)

1. Moskovskoye otdeleniye Vsesoyuznogo nauchno-issledovatel'skogo instituta galurgii.

GAZANCHIYANTS, M.G.; LASTOVTSEV, A.M.; MARTYUSHIN, I.G.; PLANOVSKIY, A.N.; KHARAKOZ, V.V.; SHNAYDER, Ye.Ye.

Apparatus for the processing of finely dispersed vegetable materials. Gidroliz. i lesokhim. prom. 18 no.6:5-6 '65. (MIRA 18:9)

1. Moskovskiy institut khimicheskogo mashinostroyeniya (for all except Shnayder). 2. Vsesoyuznyy nauchno-issledovatel'skiy institut biosinteza belkovykh veshchestv (for Shnayder).

[redacted], D.I., [redacted]

Removal of chambers for welding from water walls. Energetik 13 no.6:16-17 Je '65. (MIRA 18:7)

28(1) PHASE I BOOK EXPLOITATION SOV/1432

Shnayderman, Iosif Berkovich

Tabulyatory T-4M i T-4MI; elementy avtomaticheskogo upravleniya i metody kommutatsii (T-4M and T-4MI Tabulating Machines; Elements of Automatic Control and Methods of Switching) Moscow, Metallurgizdat, 1958. 199 p. 2,000 copies printed.

Ed.: Bulavko, Yu. M.; Ed. of Publishing House: Lanovskaya, M.R.; Tech. Ed.: Mikhaylova, V.V.

Purpose: The book is intended for design engineers dealing with automatic calculation. It may also be useful to students of vuzes and vocational schools studying computing machines.

COVERAGE: The author describes automatic control of the T-4M and T-4MI tabulators and the IP-45 summary punch. He also presents a detailed discussion of the control panel hubs and control levers and shows their interaction with the electrical circuits. The operation and switching of the various mechanisms of the tabulators are also described. No personalities are mentioned. There are no references.

Card 1/6

SHNAYDERMAN, I. YA.

Shnayderman, I. Ya. "Analytical computation of cutter profile for cutting trapezoidal high-pitch thread," Izvestiya Kiyevsk, Politekhn, in-ta, Vol VIII, 1948 (on cover:1949), p. 231-46

SO: U-5241, 17 December 1953, (Letopis 'Zhurnal 'nykh Statey, No. 26, 1949)

SHNAYDERMAN, I.Ya., kandidat tekhnicheskikh nauk.

Waviness of surfaces following threading by the vortex method.
Vest.mash. 36 no.10:35-39 O '56. (MLRA 9:11)
(Screw cutting)

ANISIMOV, Yefim Georgiyevich; SHNAYDERMAN, I.Ya., kand.tekhn.nauk, retsenzent; ONISHCHENKO, N.P., inzh., red.; RUDENSKIY, Ya.V., tekhn.red.

[Design of machine-tool attachments in lot production] Proektirovanie stanochnykh prisposobleniĭ v seriinom proizvodstve. Moskva, Gos.nauchno-tekhn.izd-vo mashinostroit.lit-ry, 1959. 165 p. (MIRA 13:1)
(Machine tools--Attachments)

25(2) SOV/117-59-4-10/36

AUTHOR: Shnayderman, I.Ya., Candidate of Technical Sciences

TITLE: A Mobile Rest for Lathes.

PERIODICAL: Mashinostroitel', 1959, Nr 4, pp 23-26 (USSR)

ABSTRACT: Design and operational information is presented on a mobile lathe rest developed by the author in two design variations (with one support jaw and manual control, and with automatic control of the jaws), both for machining non-rigid stepped shafts with increasing and decreasing step diameters or with alternating smaller and larger step diameters in a range of 60-80 mm. The first variation (Figures 1 and 2) corresponds to the dimensions of the lathe "1A62", the second (Figure 3) is for use with automated tool control (as in work with the hydro-tracer toolrest, or the V.K. Seminskiy attachment). The inventions were registered in 1957 (Author's Certificate Nr 103228 of 3 Apr 1957). There are 3 sets of diagrams.

Card 1/1

SHNAYDERMAN, I.Ya.; KIRILYUK, Yu.Ye.

Head with quick-interchangeable cutting-tool holders. Mashinostroitel' no.9:26 S '60. (MIRA 13:9)
(Lathes)

SHNAYDERMAN, I.Ya., kand.tekhn.nauk

Adjusting multicut heads in cutting screw threads by the vortex method. Vest.mash. 40 no.10:67-68 O '60. (MIRA 13:10)
(Screw cutting)

SHNAYDERMAN, I.Ya.

Graphic method for selecting cutting conditions. Mashinostroitel' no.1:24-25 Ja '61. (MIRA 14:3)
(Metal cutting)

SHNAYDERMAN, I.Ya., kand.tekhn.nauk

Development and introduction of multiple machining on turret lathes.
Mashinostroenie no.2:9-15 Mr-Ap '62. (MIRA 15:4)

1. Kiyevskiy politekhnicheskiy institut.
(Factory management) (Lathes)

SHNAYDERMAN, I.Ya., kand. tekhn. nauk

Automatic thread-rolling unit. Mashinostroitel' no.5:5-6 My '65.
(MIRA 18:5)

KARTAVOV, Sergey Alekseyevich, prof.; LEVCHENKO, Andrey Matveyevich, kand. tekhn. nauk; RUDNIK, Sergey Sergeyevich, doktor tekhn. nauk; BOVSUNOVSKIY, Yakov Ivanovich, kand. tekhn. nauk; BAZHENOV, Ivan Ivanovich, kand. tekhn. nauk; KOVALENKO, Vladimir Vladimirovich, kand. tekhn. nauk; LOMACHENKO, Zinaida Nikolayevna, kand. tekhn. nauk; MIL'SHTEYN, Mark Zel'manovich, kand. tekhn. nauk; RADCHENKO, Yuliya Gavrilovna, kand. tekhn. nauk; REZNICHENKO, Mikhail Petrovich, kand. tekhn. nauk; TRUBENOK, Aleksandr Davidovich, kand. tekhn. nauk; KHRISTICH, Zakhar Dem'yanovich, kand. tekhn. nauk; SHNAYDERMAN, Isay Yakovlevich, kand. tekhn. nauk; GOLUBOV, N.P., kand. tekhn. nauk, retsenzent; DUMANSKAYA, V.A., kand. tekhn. nauk, retsenzent; MAKSIMOV, G.D., kand. tekhn. nauk, retsenzent; YAKOVENKO, G.A., kand. tekhn. nauk, retsenzent

[Technology of the manufacture of machinery] Tekhnologiia mashinostroeniia. [By] S.A.Kartavov i dr. Kiev, Tekhnika, 1965. 526 p. (MIRA 18:7)

1. Kafedra tekhnologii mashinostroyeniya Kiyevskogo politekhnicheskogo instituta (for all except Golubov, Maksimov, Yakovenko).

SHMIDT, L. (Alma-Ata)

An important factor in the increase of material self-interest in agricultural production. Vop. ekon. no. 2:146-149 F '61. (MIRA 14:2)

(Kazakhstan--Agricultural wages)

SHNAYDERMAN, L.M.

USSR

Liver function in hypertension. T. I. Fursova and L. M. Shnaiderman (Med. Inst. Chelyabinsk). *Klin. Med.* (U.S.S.R.) 32, No. 11, 70(1954).—Slight impairment of carbohydrate metabolism and a significant disturbance of detoxifying function were noticed among a no. of patients. The Takata-Aara and Weltman tests were pos. in a majority of cases. Bilirubin was normal. A. S. Mirkin

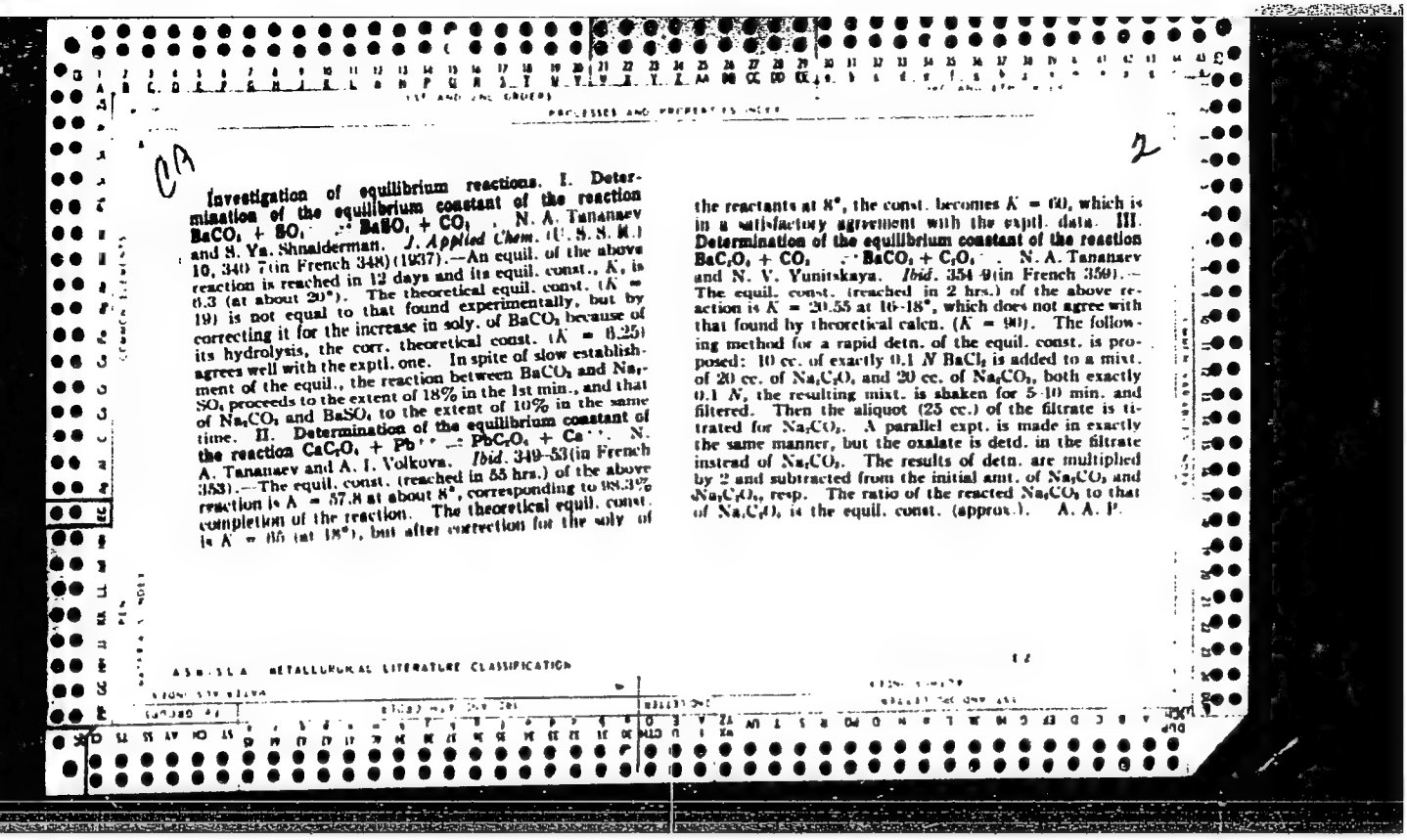
SHNAYDERMAN, P.Ya., vrach

Dental caries in children of Chistopol' in the Tatar A.S.S.R. Vop. obshchei stom. 17:7-8 '64.
(MIRA 18:11)

SHNAYDERMAN, S.Ya.; CHERNAYA, N.V.

Pyrocatechinate and pyrogallate complexes of titanium in methanol and water-methanol solutions. Zhur. neorg. khim. 10 no.1:224-230 Ja '65. (MIRA 18:11)

1. Submitted June 5, 1963.

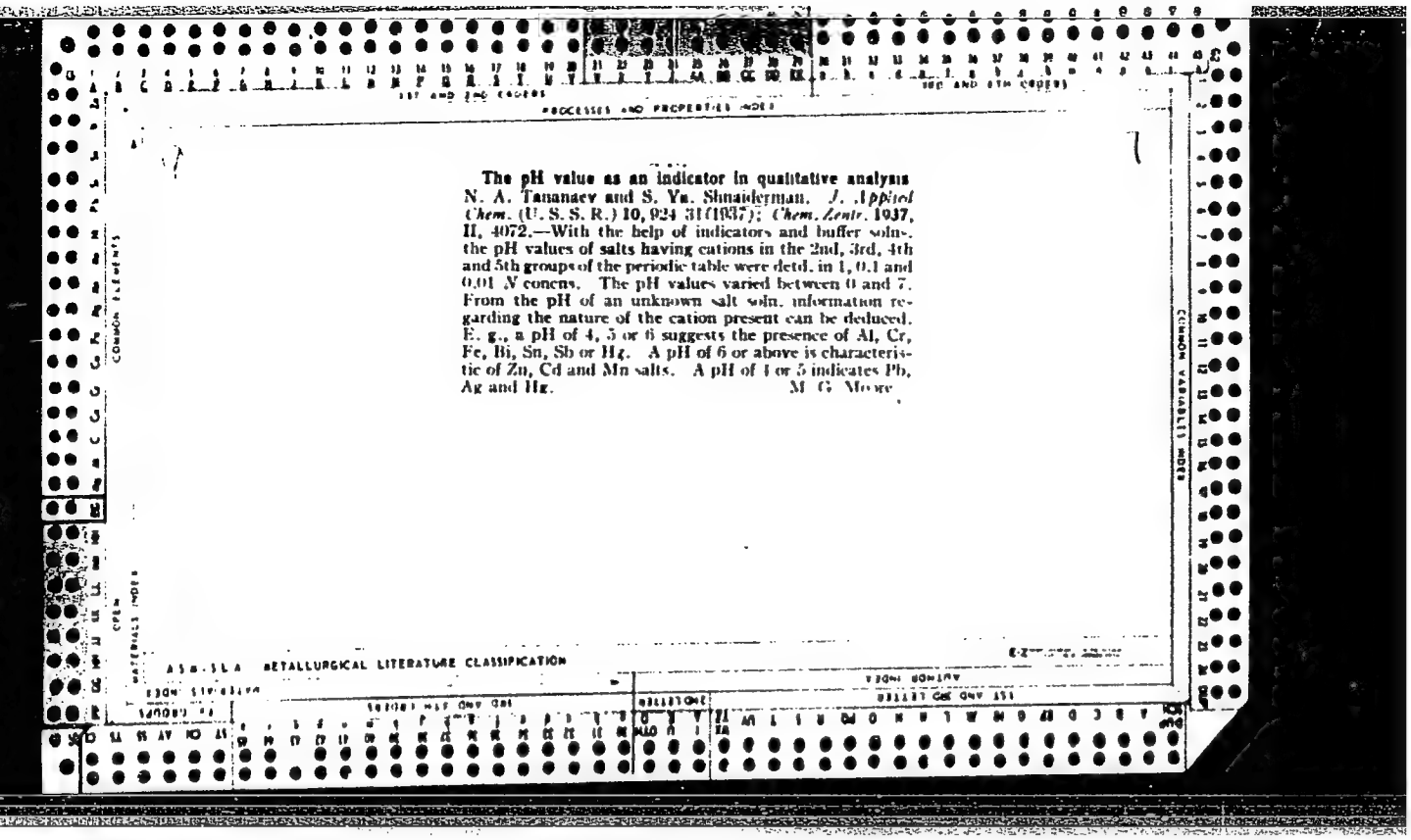


CA

2

Investigation of equilibrium reactions. I. Determination of the equilibrium constant of the reaction $BaCO_3 + SO_4^{--} \rightleftarrows BaSO_4 + CO_3^{--}$. N. A. Tananaev and S. Ya. Shnaiderman. *J. Applied Chem.* (U. S. S. R.) 10, 340-7(in French 348)(1937).—An equil. of the above reaction is reached in 12 days and its equil. const., K, is 0.3 (at about 20°). The theoretical equil. const. (K = 19) is not equal to that found experimentally, but by correcting it for the increase in soly. of $BaCO_3$ because of its hydrolysis, the corr. theoretical const. (K = 0.25) agrees well with the exptl. one. In spite of slow establishment of the equil., the reaction between $BaCO_3$ and Na_2SO_4 proceeds to the extent of 18% in the 1st min., and that of Na_2CO_3 and $BaSO_4$ to the extent of 10% in the same time. II. **Determination of the equilibrium constant of the reaction $CaC_2O_4 + Pb^{++} \rightleftarrows PbC_2O_4 + Ca^{++}$.** N. A. Tananaev and A. I. Volkova. *Ibid.* 349-53(in French 353).—The equil. const. (reached in 55 hrs.) of the above reaction is K = 57.8 at about 8°, corresponding to 88.3% completion of the reaction. The theoretical equil. const. is K = 86 (at 18°), but after correction for the soly. of the reactants at 8°, the const. becomes K = 60, which is in a satisfactory agreement with the exptl. data. III. **Determination of the equilibrium constant of the reaction $BaC_2O_4 + CO_3^{--} \rightleftarrows BaCO_3 + C_2O_4^{--}$.** N. A. Tananaev and N. V. Yunitskaya. *Ibid.* 354-9(in French 359).—The equil. const. (reached in 2 hrs.) of the above reaction is K = 20.55 at 16-18°, which does not agree with that found by theoretical calcn. (K = 90). The following method for a rapid detn. of the equil. const. is proposed: 10 cc. of exactly 0.1 N $BaCl_2$ is added to a mixt. of 20 cc. of $Na_2C_2O_4$ and 20 cc. of Na_2CO_3, both exactly 0.1 N, the resulting mixt. is shaken for 5-10 min. and filtered. Then the aliquot (25 cc.) of the filtrate is titrated for Na_2CO_3. A parallel expt. is made in exactly the same manner, but the oxalate is detd. in the filtrate instead of Na_2CO_3. The results of detn. are multiplied by 2 and subtracted from the initial amt. of Na_2CO_3 and $Na_2C_2O_4$, resp. The ratio of the reacted Na_2CO_3 to that of $Na_2C_2O_4$ is the equil. const. (approx.). A. A. P.

ASM-SLA METALLURGICAL LITERATURE CLASSIFICATION



The pH value as an indicator in qualitative analysis. N. A. Tananaev and S. Ya. Shnaiderman. *J. Applied Chem.* (U. S. S. R.) **10**, 924-31 (1937); *Chem. Zentr.* **1937**, II, 4072.—With the help of indicators and buffer solns. the pH values of salts having cations in the 2nd, 3rd, 4th and 5th groups of the periodic table were detd. in 1, 0.1 and 0.01 *N* concns. The pH values varied between 0 and 7. From the pH of an unknown salt soln. information regarding the nature of the cation present can be deduced. E. g., a pH of 4, 5 or 6 suggests the presence of Al, Cr, Fe, Bi, Sn, Sb or Hg. A pH of 6 or above is characteristic of Zn, Cd and Mn salts. A pH of 4 or 5 indicates Pb, Ag and Hg. M. G. Moore

SHNAYDERMAN, S. Ya.

Chromotropic acid as analytical reagent. I. Interaction of copper with chromotropic acid. S. Ya. Shnaiderman. *Ukrain. Khim. Zhur.* 19, 327-30(1953); *Referat. Zhur., Khim.* 1954, No. 13231.—At pH 5–11, Cu^{++} reacts with chromotropic acid to give a red soln. The optimum conditions are pH not below 9, heating to incipient boiling, and 20-fold excess of reagent. The reaction cannot be used for detg. Cu because the intensity of the color changes with time. To the analyzed soln., add an acetate-ammoniacal buffer at pH 5–11 and an excess of chromotropic acid. Heat the mixt. to boiling and cool in water. Fe, Sn, Sb, Bi, and Mn compds. which hydrolyze at this pH interval interfere and their effect is obviated by filtering off the ppt. formed before chromotropic acid is added. The sensitivity of this reaction is 3 γ Cu in 20 ml. soln. M. Hoseh

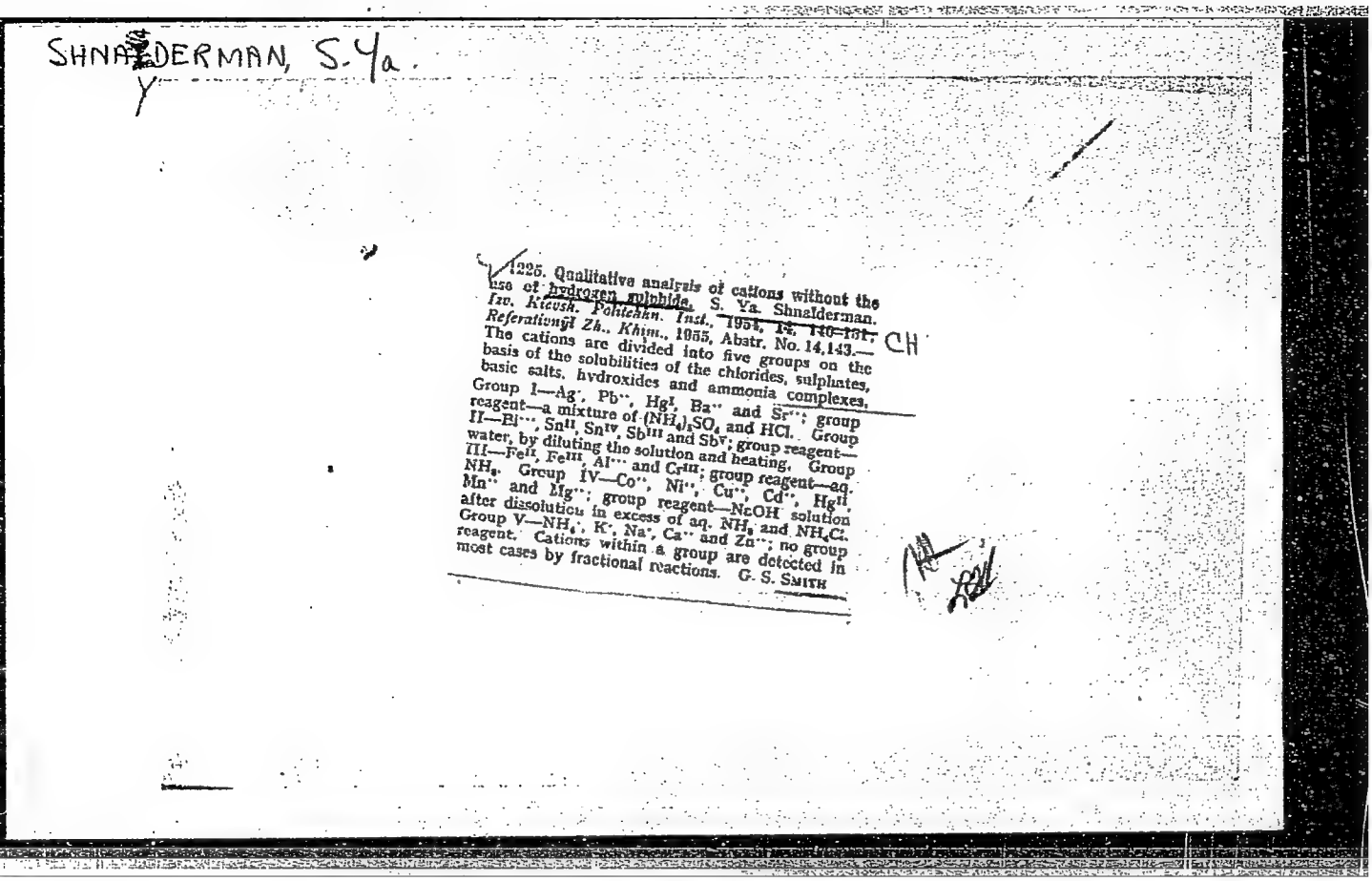
MET

SHNAYDERMAN, S.Ya. USSR

Chromotropic acid as analytical reagent. II. Reactions of chromotropic acid with some ions. S. Ya. Shnaiderman and N. P. Movchan. *Ukrain. Khim. Zhur.* 19, 429–33(1953); *Referat. Zhur., Khim.* 1954, No. 18504; cf. *C.A.* 49, 1469h.—The ions, the color, and the pH at which the color appears in reactions with chromotropic acid are: Fe^{+++}, green, 1.6–6.2; Cu^{++}, red-brown, 5.0–11.0; Hg^{++}, yellow, 3.0–6.0; Ag^{+}, bright-yellow ppt., 0–6.0 and dark-brown ppt., 8.0–10.0; Ti^{4+}, red, 0.8–3.8, orange, 4.0–5.0, yellow, 5.1–9.0; UO_2^{++}, red-brown, 4.0–10.0; CrO_4^{--}, red, $[H^+] = 10N\text{-}10^{-1}$ N; NO_2^-, yellow, $[H^+] = 10N\text{-}10^{-4}N$; WO_4^{--}, yellow, 4.3–6.0, red, 6.2–10.0; MoO_4^{--}, yellow, 4.0–7.0, red, 7.0–10.5; VO_3^-, yellow, 0–6.0, red, 8.0–10.0; NO_3^-, yellow, concd. H_2SO_4. The optical d. of the red compd. of the chromotropic acid with Ti^{4+} reaches a max. at pH 3.9 and the yellow compd. at pH 4.3–7.8. The max. optical d. for Fe^{+++} is at pH 5.0, for UO_2^{++}, WO_4^{--}, and MoO_4^{--} compds. at pH 6–9, and for NO_2^- in 1–5N H_2SO_4. Fe^{+++} interferes with the colorimetric detn. of Ti^{4+} and at pH < 2.5 also VO_3^- because Fe^{+++} compds. with chromotropic acid have an appreciable optical d. at pH 2–4 and VO_3^- at pH 0.8–2. With the detn. of CrO_4^{--} only NO_2^- interferes. Chromotropic acid does not react with NH_4, Na, K, Ca, Sr, Ba, Fe^{++}, Cr^{+++}, Ni, Co, Zn, Cd, Hg^{+}, Bi, Sn, Sb, Pb, Ti^{+++}, Be, Th, Zr, Tl, La, Ce, Nd, Pr, Rb, Cs, phosphates, $S_2O_3^{--}$, SO_3^{--}, SO_4^{--}, AsO_3^{--}, CO_3^{--}, SiO_3^{--}, $B_4O_7^{--}$, F^-, Cl^-, Br^-, I^-, CNS^-, $Fe(CN)_6^{---}$, $Fe(CN)_6^{4-}$, S^{--}, Me, COO^-, SeO_3^{--}, TeO_3^{--}, IO_3^-, BrO_3^-, and ClO_3^-. The expts. were carried out with solns. contg. 1.25×10^{-4} g.-ion/l. The chromotropic acid soln. contained 0.25×10^{-4} g. mol./l.

M. Hosch

(1)



SHNAYDERMAN, S. Ya.

1225. Qualitative analysis of cations without the use of hydrogen sulphide. S. Ya. Shnalderman. *Izv. Kievsk. Politekhn. Inst.*, 1954, 18, 140–151. *Referativnyi Zh., Khim.*, 1955, Abstr. No. 14,143.— The cations are divided into five groups on the basis of the solubilities of the chlorides, sulphates, basic salts, hydroxides and ammonia complexes. Group I—$Ag^{\cdot}$, $Pb^{\cdot\cdot}$, Hg^{I}, $Ba^{\cdot\cdot}$ and $Sr^{\cdot\cdot}$; group reagent—a mixture of $(NH_4)_2SO_4$ and HCl. Group II—$Bi^{\cdot\cdot\cdot}$, Sn^{II}, Sn^{IV}, Sb^{III} and Sb^{V}; group reagent—water, by diluting the solution and heating. Group III—Fe^{II}, Fe^{III}, $Al^{\cdot\cdot\cdot}$ and Cr^{III}; group reagent—aq. NH_3. Group IV—$Co^{\cdot\cdot}$, $Ni^{\cdot\cdot}$, $Cu^{\cdot\cdot}$, $Cd^{\cdot\cdot}$, Hg^{II}, $Mn^{\cdot\cdot}$ and $Mg^{\cdot\cdot}$; group reagent—NaOH solution after dissolution in excess of aq. NH_3 and NH_4Cl. Group V—$NH_4^{\cdot}$, $K^{\cdot}$, $Na^{\cdot}$, $Ca^{\cdot\cdot}$ and $Zn^{\cdot\cdot}$; no group reagent. Cations within a group are detected in most cases by fractional reactions. G. S. SMITH

CH

SHNAIDERMAN, S. Ya.

4

1314. Chromotropic acid as an analytical reagent. III. Interaction of nitrites with chromotropic acid. S. Ya. Shnaiderman. *Izv. Kievsk. Politekhn. Inst.*, 1955, 14, 102–106; *Referativnyĭ Zh., Khim.*, 1956, Abstr. No. 14,251.—Chromotropic acid gives with solutions of nitrites a yellow colour, the intensity of which depends on the acidity and time after mixing. The mol. concn. of the reagent should be 5 times that of the nitrite. The concn. of acid, H_2SO_4 or HCl, should be between N and 5 N. Beer's law is not obeyed. To detect NO_2', 1 to 2 ml of the solution are mixed with 1 ml of 0·01 M reagent solution and 1 to 2 ml of 3 to 5 N mineral acid. In the presence of NO_2' a yellow colour appears after 5 min. and its intensity increases with time. Ferro- and ferri-cyanides, CrO_4'' and Cr_2O_7'' interfere, but NO_3' do not. As little as 5 μg of NO_2' can be detected in 50 ml viewed through a 100-mm depth. To determine NO_2', 25 ml of solution containing 0·05 to 0·2 mg are treated with 2 to 3 ml of 0·01 M reagent and 10 ml of 5 N H_2SO_4, the solution is diluted to 50 ml and the extinction is measured at 400 to 500 mμ with a blue filter after 2 hr. A calibration curve is obtained under the same conditions.

G. S. Smith

CH

SHNAYDERMAN, S. YA.

USSR/ Chemistry - Analytical chemistry

Card 1/1 Pub. 116 - 19/25

Authors : Shnayderman, S. Ya.

Title : Phenols as nitrite reagents

Periodical : Ukr. khim. zhur. 21/1, 99-103, 1955

Abstract : It was established esperimentally that phenol and naphthol solutions can be utilized in the role of analytical reagents for the detection of small amounts of nitrites. The reaction sensitivity varies between 0.5 and 5mkg/ A close relation was established between the color intensity and the acidity of the solution, surplus of the reagent, time of reaction and nitrite concentration. A method for colorimetric determination of nitrites was also introduced. Three USSR references (1946-1954). Table; graphs.

Institution: The Polytechnicum, Kiev

Submitted : December 28, 1953 and June 10, 1954

Shnayderman, S. Ya.

USSR/ Chemistry - Analytical chemistry

Card 1/1 Pub. 116 - 21/24

Authors : Shnayderman, S. Ya.

Title : Discovery and colorimetric determination of Au with the aid of ascorbic acid

Periodical : Ukr. khim. zhur. 21/2, 261-264, 1955

Abstract : It was proven experimentally that ascorbic acid is a highly sensitive reagent for Au. The acid was found to be the most suitable medium for qualitative discovery and colorimetric determination of Au. The optimum conditions favorable for carrying out the reaction at low acidity were established at 3 - 6 pH values. Data regarding the optical density and transparency of colloidal Au solutions are included. Three references: 2 USSR and 1 English (1927-1942). Tables; graphs.

Institution : The Kiev Polytechnic Inst.

Submitted : February 20, 1954

SHNAYDERMAN, S. YA.

USSR/Analytical Chemistry - General Questions G-1

Abs Jour : Referat Zhur - Khimiya, No 3, 1957, 8365

Author : Shnayderman, S. Ya.
Inst : Kiev Polytechnic Institute
Title : Chromotropic Acid as an Analytical Reagent. Communication IV. Determination of Titanium in the Presence of Iron and Vanadium

Orig Pub : Izv. Kievsk. politekhn. in-ta, 1956, Vol 17, 197-203

Abstract : Cast iron or steel (0.5 gms) are dissolved in 20 ml H_2SO_4 (1 : 3); when the samples have dissolved completely, 1-2 ml conc. HNO_3 are added to destroy the carbides and oxidize the iron. The solution is cooled and transferred to a 50 ml volumetric flask; water is added up to the mark, 1-2 ml of the solution (depending on the expected Ti content) are transferred to a 50-ml volumetric flask with a micropipette, and carefully neutralized with 10% ammonia until a weakly acidic reaction is obtained, characterized by the appearance of slowly dissolving turbidity, at which point 1-2 ml of 3% ascorbic acid are added. The contents of the flask are shaken, 1-2 ml of 2% chromotropic acid are added and the solu-

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USSR/Analytical Chemistry - General Questions G-1

Abs Jour : Referat Zhur - Khimiya, No 3, 1957, 8365

tion is diluted to the mark with a pH 2-3.5 buffer solution; the intensity of the color of the solution is measured. At pH 4.3-7.8 the yellowish Ti complex exhibits maximum constant optical density. The discoloration of the solution by sodium fluoride stops at pH 4.7. For communication 3 see RZhKhim, 1955, 14251.

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-5-

Abs Jour : Referat Zhur - Khimiya, No 3, 1957, 8433

the solution is adjusted to the mark, the solution is stirred, and its optical density is determined with a photocolorimeter. The Au content of the solution is determined from a calibration curve.

SHNAYDERMAN, S. Ya.

Complex formation in the stystem titanium - chromotropic acid. (MIRA 10:12)
Zhur.neorg.khim. 2 no. 9:2122-2125 S '57
(Titanium) (Naphthalenedisulfonic acid)

SHNAYDERMAN, S.Yn., dots., kand.khim.nauk; ROBEROVA, I.B.

Relationship of phenols to ion series. Izv. KPI 20:108-126 '57.
(Phenols) (Ions) (Colorimetry) (MIRA 11:3)

AUTHOR: Shnayderman, S. Ya. 73-1-18/26

TITLE: Colour Reactions of Titanium with Phenols. (Tsvetnyye Reaktsii Titana S Fenolami.)

PERIODICAL: Ukrainskiy Khimicheskiy Zhurnal, 1957, Vol. 23, No.1, pp 92 - 96 (USSR).

ABSTRACT: Titanium gives colour reactions with many phenols. The colour depends on the character of the phenols, the concentration of titanium and the acidity of the solution. Particularly strong tendencies to complex-formation occurs in the case of poly-atomic phenols containing the -OH or -OH and -COOH groups occur side by side. Experiments on the inter-reaction of iron-, molybdenum-, vanadium, titanium- and cerium-salts with various phenols were carried out. On the interaction of some phenols and hydroxy-acids with diluted solutions of the above salts (of the order 10^{-2} - 10^{-4} mole) an intensive coloration of the solutions is only obtained when pyrogallol, pytocatechol and gallic acid are used as reagents. Strong phenols either do not react at all or form slightly coloured compounds with the salts of iron and titanium. Experimental data on the reaction of titanium with phenols in strong and weak acid solutions are discussed. The dependance of colour on the concentration of the

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Colour Reactions of Titanium with Phenols. 73-1-18/26

sulphuric acid, on the excess reagent and on time was investigated spectrophotometrically. Diagrams giving the composition of Ti-chromotropic substances at pH 3-4, the relation of the optical density and the molar concentrations of phenols and titanium (graph 2), on the percentage content of sulphuric acid for various phenols (graphs 3 and 4) illustrate results obtained during spectrophotometric investigations. There are 5 diagrams and 5 references, 2 of which are Slavic.

SUBMITTED: August, 21, 1956.

ASSOCIATION: Kiyev Polytechnical Institute. (Kiyevskiy Polytekhnicheskiy Institut.)

AVAILABLE: Library of Congress

Card 2/2

AUTHORS: Shnayderman, S. Ya., Khrustalev, G. I. (Deceased)

SOV/79-29-1-6/74

TITLE: The Reaction of Molybdate and Vanadate With Phenols in Aqueous Solutions and Concentrated Sulfuric Acid (Vzaimodeystviye molibdata i vanadata s fenolami v vodnykh rastvorakh i v kontsentrirovannoy sernoy kislote)

PERIODICAL: Zhurnal obshchey khimii, 1959, Vol 29, Nr 1, pp 20-27 (USSR)

ABSTRACT: Vanadium and molybdenum with some phenols form colored compounds in aqueous solutions and in concentrated sulfuric acid. Individual phenols are used as reagents in the case of colorimetric determination of vanadate and molybdate (Refs 1,2). Apart from the investigations carried out by Levy (Ref 3) where the colors of the solutions in concentrated sulfuric acid are given, the authors found no mentions in publications concerning problems of the reaction of vanadium (5 +) and molybdenum (6 +) with oxy-compounds. This is the reason for their dealing with the problem. Experiments yielded the following results: Vanadates and molybdates yield intensely colored solutions with oxy-compounds at a percentage of 4-10% only in the case if their complex forming groups are

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in ortho-position towards one another. In concentrated sulfuric solution oxy-compounds change their color independently of their structure. The color of the concentration depends to a great extent on the surplus of the reagent and in highly acid solutions on the concentration of sulfuric acid. The most intense color is caused by sulfuric acid. In case the acid is diluted it decreases. It is not easy to find a range for the concentration of sulfuric acid in which the color is stabilized. The dependence between the color of the solutions upon the molar ration of the oxy-compound and vanadate or molybdate points to the possibility of a formation of complex compounds. Thus, the character of the reactions of vanadate and molybdate was determined with phenols in a weakly acid medium, and the composition of the compounds in concentrated sulfuric acid. The spectrophotometric investigation, the dependence of the color of the concentration of sulfuric acid on the excess of reagent and time are given by 8 diagrams. There are 8 figures, 4 tables, and 11 references, 8 of which are Soviet.

The Reaction of Molybdate and Vanadate With Phenols in Aqueous Solutions and Concentrated Sulfuric Acid SOV/79-29-1-o/74

ASSOCIATION: Kiyevskiy politekhnicheskiy institut (Kiyev Polytechnical Institute)

SUBMITTED: August 4, 1957

Card 3/3

S/153/60/003/02/08/034
B011/B003

5.5300

AUTHOR: Shnayderman, S. Ya.

TITLE: Photometric Investigation of the Reaction of Titanium (IV) With Phenols in Acetic Acid

PERIODICAL: Izvestiya vysshikh uchebnykh zavedeniy. Khimiya i khimicheskaya tekhnologiya, 1960, Vol. 3, No. 2, pp. 258-264

TEXT: In the article under review, the author intended to investigate the complex formation of titanium with phenols in acetic medium. Furthermore, the most important physicochemical characteristics of the compounds formed and the possibilities of using these substances for the colorimetric determination of titanium were to be studied. The results obtained by the author with respect to the system titanium salt - phenol - acetic acid are listed. The author mixed titanium solutions with phenols which were used in optimum excess so that the titanium concentration was 10^{-3} g-ion/l, whereas the total volume of the

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Photometric Investigation of the Reaction of Titanium (IV) With Phenols in Acetic Acid

S/153/60/003/02/08/034
B011/B003

solution was 1 ml. These mixtures were diluted to 10 ml with glacial acetic acid. After 5 minutes the color of the solution was visually ascertained. The sensitivity of the reactions in concentrated acetic acid is shown in Table (p. 259). Hence it may be seen that a strong coloring occurs which may be used for the colorimetric analysis. The most sensitive reagent is chromotropic acid. Other phenols and their derivatives were the following: phenol, guaiacol, thymol, pyrocatechol, resorcinol, hydroquinone, pyrogallol, phloroglucinol, gallic acid, α- and β-naphthol. The author proved that all hydroxy compounds form complexes with titanium (IV) in acetic acid only when the acetic acid is concentrated. An exception is made by the complex of the titanium chromotropic acid and titanium gallic acid. The titanium complex with the chromotropic acid is destroyed in concentrated acetic acid. The author proved that titanium can be analytically determined with chromotropic acid in a wide concentration range of the acetic acid. The composition of the complex of the titanium chromotropic acid in a 25 - 75% acid in a ratio of titanium : chromotropic acid = 1 : 4

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Photometric Investigation of the Reaction of Titanium (IV) With Phenols in Acetic Acid

S/153/60/003/02/08/034
B011/B003

was determined by means of the method of isomolar series. The reaction of its formation is represented by the author as follows: $Ti^{4+} + 4\left[C_{10}H_4(OH)_2(SO_3)_2\right]^{2-} = Ti\left[C_{10}H_4O_2H(SO_3)_2\right]_4^{8-} + 4H^+$. A structural formula is also given. The dependence of the optical density of the phenols enumerated and their derivatives on the molar concentrations of the phenols and of titanium (Fig. 1), on the concentration of the acetic acid (Fig. 2), and on time (Fig. 3) is given in Figs. 1-3. The light-absorption curves of the titanium-chromotropic acid complex are illustrated in Fig. 4. The stable composition of this complex between 1 - 120 h is proven in Fig. 5. There are 5 figures, 1 table, and 21 references, 10 of which are Soviet.

ASSOCIATION: Kiyevskiy politekhnicheskiy institut; Kafedra analiticheskoy khimii (Kiyev Polytechnic Institute; Chair of Analytical Chemistry)

SUBMITTED: June 17, 1958

Card 3/3

SHIAY DERMAN, S.Ya.

Color reactions of titanium with phenols in strongly acidic solutions. Trudy kom. anal. khim. 11:273-284 '60. (MIRA 13:10)
(Titanium) (Phenols)

S/073/60/026/005/013/019
B004/B063

AUTHOR: Shnayderman, S. Ya.

TITLE: Phenol Complexes of Tetravalent Vanadium in Sulfuric Acid

PERIODICAL: Ukrainskiy khimicheskiy zhurnal, 1960, Vol. 26, No. 5, pp. 653 - 657

TEXT: Proceeding from Ref.1 the authors have studied the formation of colored complexes from vanadyl chloride ($VOCl_2$) and various phenols in H_2SO_4. Whereas phenol, thymol, resorcinol, hydroquinone, phloroglucinol, gallic and chromotropic acids showed no color reactions, colors were obtained from pyrocatechol, pyrogallol, and guaiacol, which were green at a concentration ratio of [phenol] : [vanadin] > 1 and brown at [phenol] : [vanadin] < 1. No color reactions occurred in hydrochloric, acetic, and phosphoric acid solutions. The colored solutions showed no Tyndall effect. The sensitivity of the color reactions to vanadium amounted to 1.9 γ/ml with guaiacol, 1.7 γ/ml with pyrocatechol, and 1.4 γ/ml with pyrogallol. This reaction was affected by Fe^{2+}, Fe^{3+},

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Phenol Complexes of Tetravalent Vanadium in Sulfuric Acid

S/073/60/026/005/013/019
B004/B063

and Cu^{2+}, and to a small degree by Cr^{3+}, and also by alkaline-earth ions when using guaiacol. Disturbing anions were NO_2^-, I^-, $S_2O_3^{2-}$, and $[Fe(CN)_6]^{3-}$. Light absorption was recorded by an CΦ-5 (SF-5) spectrophotometer, and its dependence on the concentration of H_2SO_4 was studied. At a constant sulfuric acid concentration the solutions obeyed the Beer law. In the case of pyrocatechol and guaiacol, the maximum of absorption was reached at a concentration ratio of V : phenol = 2 : 1, while it was reached at 3 : 1 in the case of pyrogallol. There are 4 figures and 2 Soviet references.

ASSOCIATION: Kiyevskiy politekhnicheskiy institut (Kiyev Polytechnic Institute)

SUBMITTED: June 10, 1958

Card 2/2

SHNAYDERMAN, S.Ya.; KALINICHENKO, I.Ye.

Complexes of titanium with pyrogallol. Izv.vys.ucheb.zav.; khim.i khim.tekh. 4 no.6:897-904 '61. (MIRA 15:3)

1. Kiyevskiy politekhnicheskiy institut, kafedra analiticheskoy khimii.

(Titanium compounds) (Pyrogallol)

SHNAYDERMAN, S.Ya.; KALINICHENKO, I.Ye.

Pyrocatechol complexes of titanium. Zhur.neorg.khim. 6 no.8:1843-1849 Ag '61. (MIRA 14:8)

(Titanium compounds) (Pyrocatechol)

SHNAYDERMAN, S.Ya.; KALINICHENKO, I.Ye.

Extraction of phenolic complexes of titanium. Ukr.khim.zhur. 27 no.3:402-407 '61. (MIRA 14:11)

1. Kiyevskiy politekhnicheskiy institut.
(Titanium compounds)
(Phenol)

SHNAYDERMAN, S.Ya.; KALINICHENKO, I.Ye.

Acetate complexes of titanium. Dokl. AN SSSR 139 no.4:910-912 Ag '61. (MIRA 14:7)

1. Kiyevskiy politekhnicheskiy institut. Predstavleno akademikom A.A. Grinbergom.

(Titanium compounds)

SHNAYDERMAN, S.Ya.; GALINKER, E.V.

Pyrocatechol complexes of uranyl. Zhur.neorg.khim. 7 no.2:279-283 F '62. (MIRA 15:3)
(Uranyl compounds) (Pyrocatechol)

SHNAYDERMAN, S.Ya.; SHERSTYUK, V.P.

Chromotropic complexes of molybdenum. Zhur.neorg.khim. 8 no.2: 457-463 F '63. (MIRA 16:5)
(Molybdenum compounds) (Naphthalenedisulfonic acid)

SHNAYDERMAN, S.Ya.

Photometric study of the system Vanadium (IV) - phenols. Zhur.-neorg.khim. 8 no.2:464-473 F '63. (MIRA 16:5)

1. Kiyevskiy politekhnicheskiy institut.
(Vanadium compounds) (Phenols)

[illegible], L.A.; [illegible], N.I.

Determination of titanium by means of chromotropic acid in [illegible] acid solutions. Izv. vys. ucheb. zav.; khim. i khim. tekh. 7 no.3:513-517 '64. (MIRA 17:10)

1. Nizhegorodskiy politekhnicheskiy institut, kafedra analiticheskoy khimii.

SHNAYDERMAN, S.Ya.; KNYAZEVA, Ye.N.

Complex formation in the systems titanium (IV) - pyrocatechol-antipyrine and titanium (IV) - pyrogallol - antipyrine. Ukr.khim. zhur. 30 no.11:1135-1141 '64. (MIRA 18:2)

1. Kiyevskiy politekhnicheskiy institut.

SHNAYDERMAN, S.Ya.; KNYAZEVA, Ye.N.

Complex formation in the systems titanium (IV) - gallic acid - antipyrine and titanium (IV) - pyrogallolcarboxylic acid - antipyrine. Ukr. khim. zhur. 31 no.1:27-32 '65. (MIRA 18:5)

1. Kiyevskiy politekhnicheskiy institut.

SHNAYDERMAN, S.Ya.; CHERNAYA, N.V.

Effect of alcohol and water-alcohol solvents on the stability of titanium phenol complexes. Zhur.neorg.khim. 11 no.1:134-137 Ja '66. (MIRA 19:1)

1. Submitted June 30, 1964.

KNYAZEVA, Ye.N.; SHNAYDERMAN, S.Ya.

Complex formation in the system titanium (IV)-o-polyphenols-pyramidon. Zhur.neorg.khim. 10 no.8:1848-1852 Ag '65. (MIRA 19:1)

1. Submitted September 16, 1964.

L 23621-66 EWT(l)/EWT(m)/EWA(d)/T/EWP(t) IJP(c) JD/JM

ACC NR: AP6009516 SOURCE CODE: UR/0413/66/000/005/0037/0037

AUTHOR: Balashov, A. N.; Shnayderman, V. I.

ORG: none

TITLE: A method for making electrical resistors from microwire. Class 21, No. 179367

SOURCE: Izobreteniya, promyshlennyye obraztsy, tovarnyye znaki, no. 5, 1966, 37

TOPIC TAGS: microwire, resistor

ABSTRACT: This Author's Certificate introduces a method for making electrical resistors from microwire in glass insulation on the basis of Author's Certificate No 114718. The metal is heated to the melting point so that the forces of surface tension convert the wire into a system of small spheres connected by thin bridges. The metal is preheated to a temperature close to the melting point and a current pulse is then sent through the wire. This pulse gives sufficient extra current to melt the wire and has a duration somewhat shorter than the time required for sphere formation.

SUB CODE: 09/ SUBM DATE: 23Dec64/ ORIG REF: 000/ OTH REF: 000

UDC: 621.316.842-
-181.4

Card 1/1

S/599/62/000/031/005/006
A066/A126

AUTHOR: Shnaydman, A.V.

TITLE: Determination of the characteristics of a constant turbulence in some layers of the free atmosphere

SOURCE: Kiyev. Ukrainskiy nauchno-issledovatel'skiy gidrometeorologicheskiy institut. Trudy, no. 31, 1962. Voprosy fiziki atmosfery, 54 - 59

TEXT: A theory is put forward for calculating the coefficient of turbulence in layers of the free atmosphere with great vertical gradients of temperature and wind velocity. The theory is based on the following assumptions: 1) The coefficient of turbulence depends only slightly on the coordinates, and the value $\bar{K}$ averaged over the layer can be used with a fairly high degree of accuracy. 2) The conversion of turbulence energy into heat may be neglected. 3) The direction of the geostrophic wind varies only slightly with altitude. Thus, the equations of motion read as follows:

$$\bar{K} \frac{d^2 u}{dz^2} + 2 \omega_z v = 0 , \qquad (1)$$

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Determination of the characteristics of a ....

$$\overline{k}\,\frac{d^2v}{dz^2} - 2\omega_z\,(u - u_g) = 0\,, \tag{2}$$

where $2\omega_z$ denotes the Coriolis parameter, and u_g is the velocity of the geostrophic wind. The equilibrium of turbulence energy is given by

$$\int_{-H}^{H} \left[\left(\frac{du}{dz}\right)^2 + \left(\frac{dv}{dz}\right)^2\right] dz = \int_{-H}^{H} \frac{g}{T}\,(\gamma_a - \gamma)\,dz\,, \tag{3}$$

and the thickness 2H of the turbulized layer is obtained from

$$\int_{-H}^{H} \left[\left(\frac{du}{dz}\right)^2 + \left(\frac{dv}{dz}\right)^2\right] dz = (1 - \mu) \int_{-\infty}^{\infty} \left[\left(\frac{du}{dz}\right)^2 + \left(\frac{dv}{dz}\right)^2\right] dz\,. \tag{4}$$

If the profile of the geostrophic wind is known, the equation of motion can be solved for three layers: $z \leq -h$; $-h \leq z \leq h$; $z \geq h$. The extensive solutions indicate that the wind may deviate from the geostrophic state both negatively and positively. This depends on the variations of the geostrophic wind with al-

Determination of the characteristics of a ....

titude and with the position of the level for which these deviations are calculated. $\bar{k}$ is obtained from the equilibrium equation for the energy of turbulence as follows:

$$\bar{k} = \frac{(\Delta u)^2 \omega_z}{0.695\, g\, \bar{\beta}}, \qquad (14)$$

where $\bar{\beta} = \frac{1}{T(\gamma_a - \bar{\gamma})}$; the thickness of the turbulized layer is

$$2H = \frac{1.39}{\sqrt{\omega_z}} \sqrt{\bar{k}}.$$

It is concluded that the turbulence in the front zone increases with increasing sharpness of the front in the temperature field, with increasing inclination of the front, and with decreasing stability of the air in the separated air masses. The coefficient of turbulence is considerable in the vicinity of the front, and the degree of turbulence must be taken into account when calculating vertical motions, the cloudiness, and similar phenomena. There are 2 figures and 2 tables.

Card 3/3

NESIS, A.I.; KIN, A.I.; SHMYRALIN, L.K.; ENNS, F.G.

X-ray and pathomorphological comparisons between cardiac changes in anthracosilicosis. Izv. AN Kazakh. SSR. Ser. med. nauk 11 no. 2:50-55 '64. (MIRA 17:7)

NESIS, A.I.; VINARIK, E.M.; DVOYRIN, V.L.; DZHANGOZINA, D.M.; KLYATSKINA, I.Ye.; FADEYEVA, Ye.I.; SHNAYDMAN, I.M.; IVAKINA, T.P.

Regression of experimental silicosis under the influence of hydrocortisone. Izv. AN Kazakh. SSR Ser. med. nauk 11 no.3: 44-49 '64 (MIRA 18:1)

SHNAYDMAN, I.M. (Karaganda)

Histochemistry of some enzymes of the energy metabolism in the process of the formation of silicotic connective tissue. Arkh. pat. 27 no.11:34-40 '65. (MIRA 18:12)

1. Kliniko-rentgenologicheskiy otdel (zav. - dotsent A.I. Nesis) Kazakhskogo nauchno-issledovatel'skogo instituta gigiyeny truda i profzabolevaniy (direktor- kand.med.nauk Z.K.Tulegenov) i kabinet gistokhimii (zav. - kand.med.nauk N.T.Raykhlin) otdela patologicheskoy anatomii opukholey cheloveka (zav. - deystvitel'nyy chlen AMN SSSR prof. N.A. Krayevskiy) Instituta eksperimental'noy i klinicheskoy onkologii (direktor - deystvitel'nyy chlen AMN SSSR prof. N.N. Blokhin) AMN SSSR. Submitted February 14, 1964.

RAYKHLIN, N.T.; SHNAYDMAN, I.M.

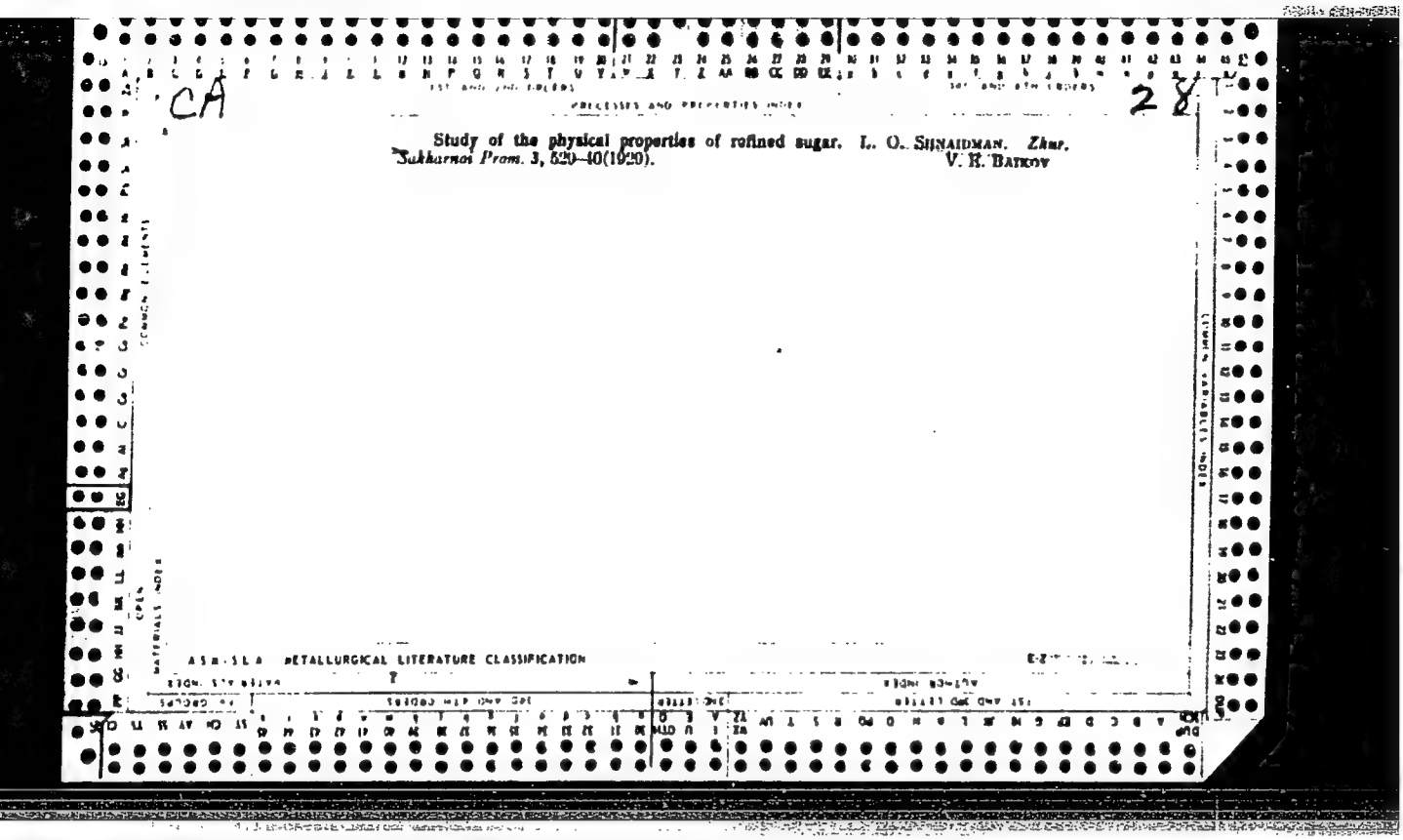
Histochemical study of the oxidation-reduction enzymes in experimental silicosis. Biul. eksp. biol. i med. 60 no. 10:112-116 O '65 (MIRA 19:1)

1. Kabinet gistokhimii (zav. - kand. med. nauk N.T. Raykhlin) otdela patologicheskoy anatomii opukholey cheloveka (zav. - deystvitel'nyy chlen AMN SSSR prof. N.A. Krayevskiy) nstituta eksperimental'noy i klinicheskoy onkologii (direktor - deystvitel'nyy chlen AMN SSSR prof. N.N. Blokhin) AMN SSSR i kliniko-rentgenologicheskiy otdel (zav. - dotsent A.I. Nesis) Nauchno-issledovatel'skogo instituta gigiyeny truda i professional'nykh zabolevaniy (direktor - kand. med. nauk Z.K. Tulegenov), Karaganda. Submitted April 10, 1964.

SHNAYDMAN, L.M., gornyy inzh.

Experiment of gas removal in the development workings by means of barrier boreholes in Mine No.26 of Karagandaugol' Combine. Ugol' 37 no.1:55-57 Ja '62. (MIRA 15:2)

1. Shakhta No.26 kombinata Karagandaugol'.
(Karaganda Basin—Mine gases)



CA

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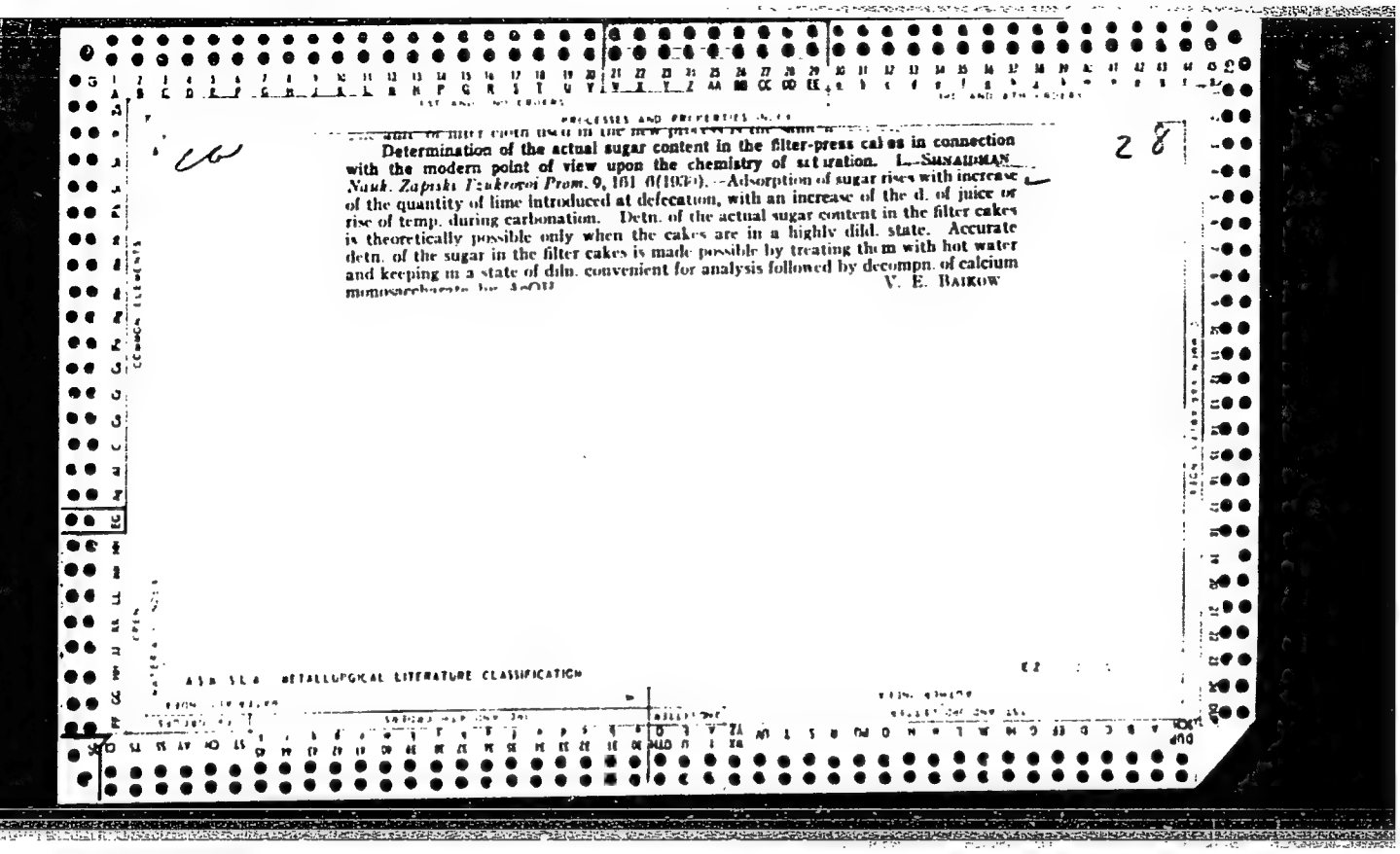
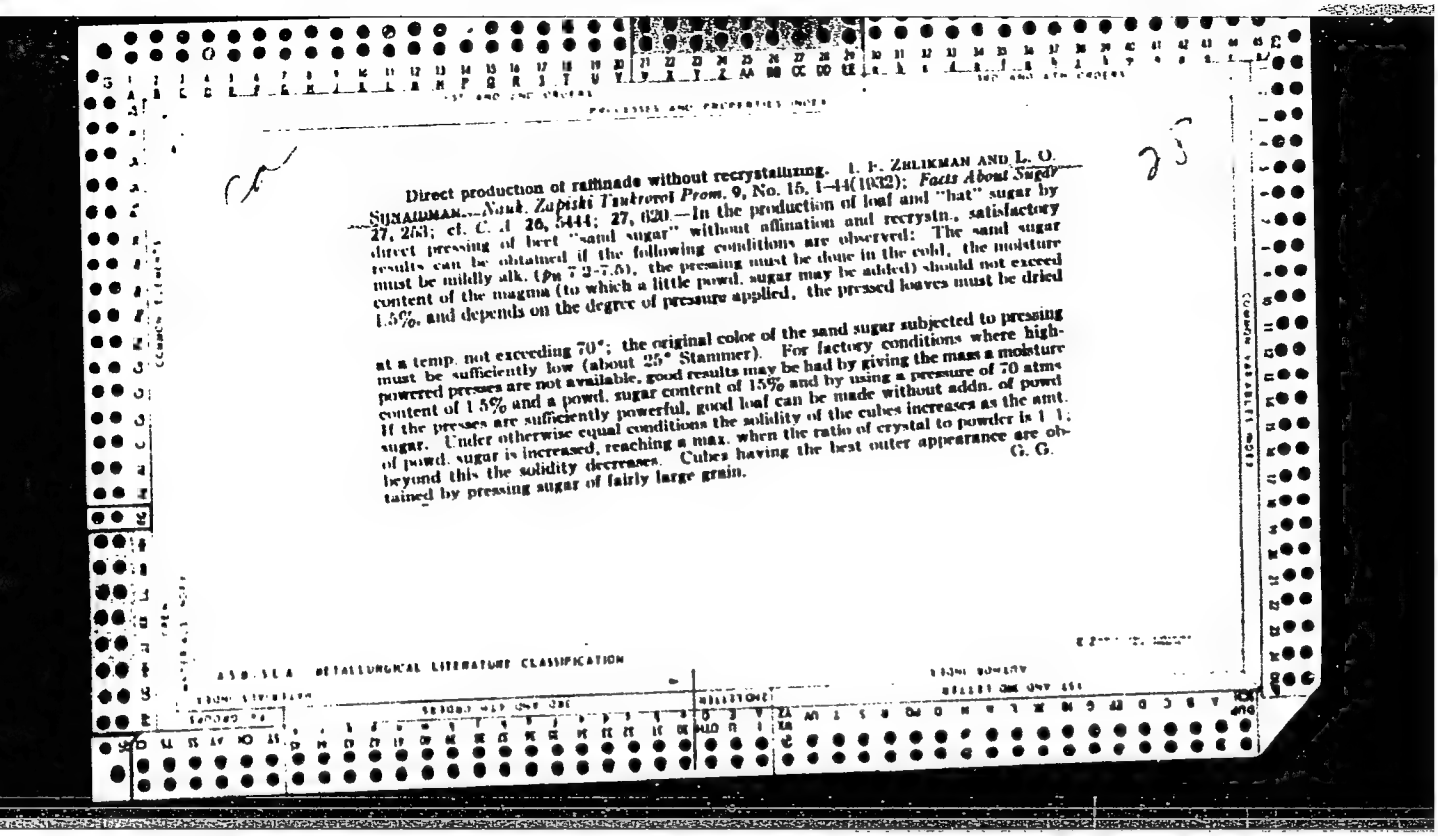
Study of the physical properties of refined sugar. L. O. SHNAIDMAN. *Zhur. Sakharnoi Prom.* 3, 529–40(1929). V. K. BATKOV

ASM-SLA METALLURGICAL LITERATURE CLASSIFICATION



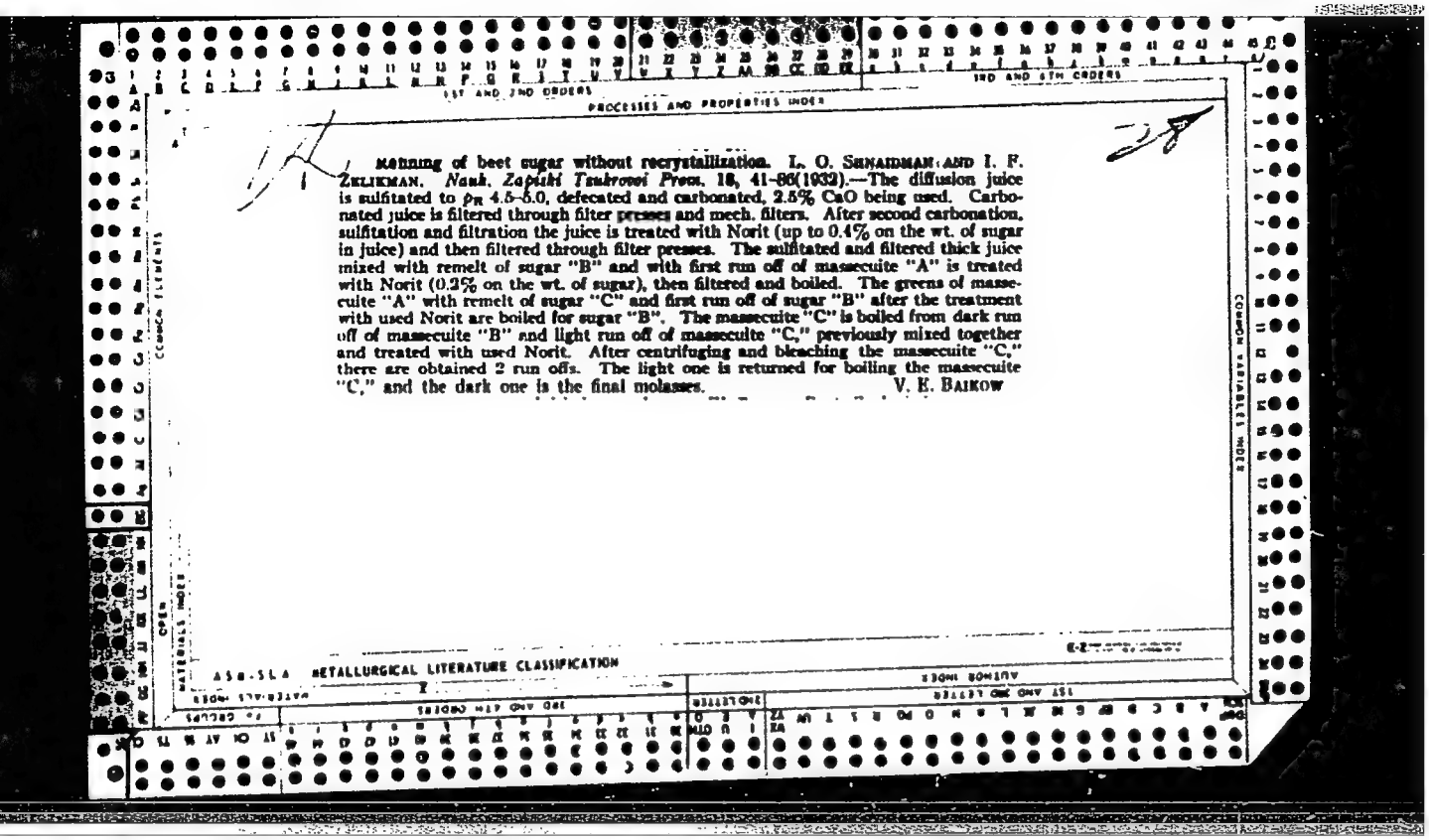
Determination of the actual sugar content in the filter-press cakes in connection with the modern point of view upon the chemistry of saturation. L. Shnaiderman. *Nauk. Zapiski Tzukrovoi Prom.* 9, 161-6(1939).—Adsorption of sugar rises with increase of the quantity of lime introduced at defecation, with an increase of the d. of juice or rise of temp. during carbonation. Detn. of the actual sugar content in the filter cakes is theoretically possible only when the cakes are in a highly dild. state. Accurate detn. of the sugar in the filter cakes is made possible by treating them with hot water and keeping in a state of dilu. convenient for analysis followed by decompn. of calcium monosaccharate by AcOH. V. E. Baikow

A53 SL6 METALLURGICAL LITERATURE CLASSIFICATION



Direct production of raffinade without recrystallizing. I. F. ZELIKMAN AND L. O. SUNAIMAN.—*Nauk. Zapiski Tsukrovoi Prom.* 9, No. 15, 1–44(1932); *Facts About Sugar* 27, 253; cf. C. A. 26, 5444; 27, 620.—In the production of loaf and "hat" sugar by direct pressing of beet "sand sugar" without affination and recrystn., satisfactory results can be obtained if the following conditions are observed: The sand sugar must be mildly alk. (*Ph* 7.2–7.6), the pressing must be done in the cold, the moisture content of the magma (to which a little powd. sugar may be added) should not exceed 1.5%, and depends on the degree of pressure applied, the pressed loaves must be dried at a temp. not exceeding 70°; the original color of the sand sugar subjected to pressing must be sufficiently low (about 25° Stammer). For factory conditions where high-powered presses are not available, good results may be had by giving the mass a moisture content of 1.5% and a powd. sugar content of 15% and by using a pressure of 70 atms. If the presses are sufficiently powerful, good loaf can be made without addn. of powd. sugar. Under otherwise equal conditions the solidity of the cubes increases as the amt. of powd. sugar is increased, reaching a max. when the ratio of crystal to powder is 1:1; beyond this the solidity decreases. Cubes having the best outer appearance are obtained by pressing sugar of fairly large grain.

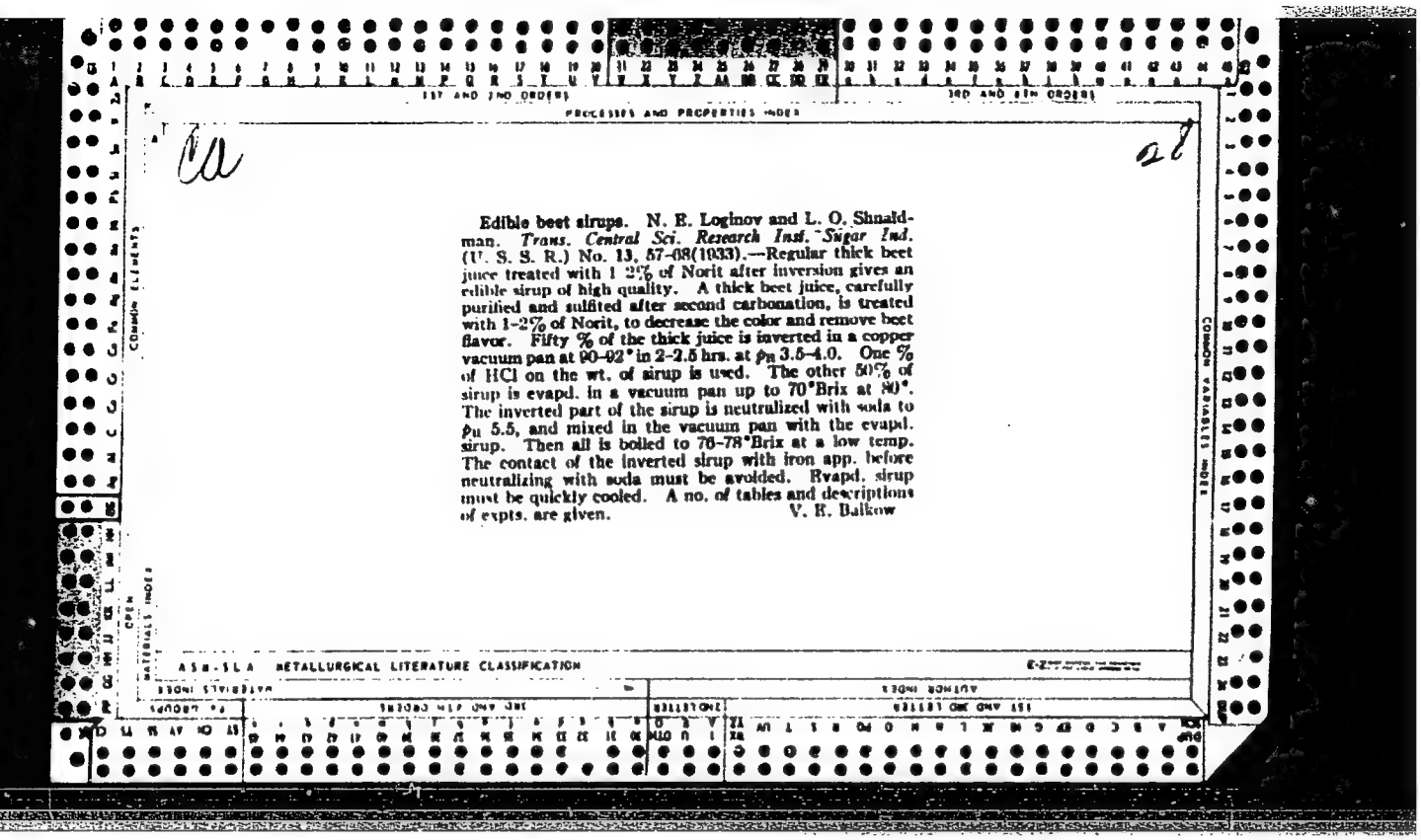
G. G.



PROCESSES AND PROPERTIES INDEX

Refining of beet sugar without recrystallization. L. O. Shnaidman and I. F. Zelikman. *Nauk. Zapiski Tsukrovoi Prom.* **18**, 41–66(1932).—The diffusion juice is sulfitated to p_H 4.5–5.0, defecated and carbonated, 2.5% CaO being used. Carbonated juice is filtered through filter presses and mech. filters. After second carbonation, sulfitation and filtration the juice is treated with Norit (up to 0.4% on the wt. of sugar in juice) and then filtered through filter presses. The sulfitated and filtered thick juice mixed with remelt of sugar "B" and with first run off of massecuite "A" is treated with Norit (0.2% on the wt. of sugar), then filtered and boiled. The greens of massecuite "A" with remelt of sugar "C" and first run off of sugar "B" after the treatment with used Norit are boiled for sugar "B". The massecuite "C" is boiled from dark run off of massecuite "B" and light run off of massecuite "C," previously mixed together and treated with used Norit. After centrifuging and bleaching the massecuite "C," there are obtained 2 run offs. The light one is returned for boiling the massecuite "C," and the dark one is the final molasses. V. E. Baikow

METALLURGICAL LITERATURE CLASSIFICATION



Ca

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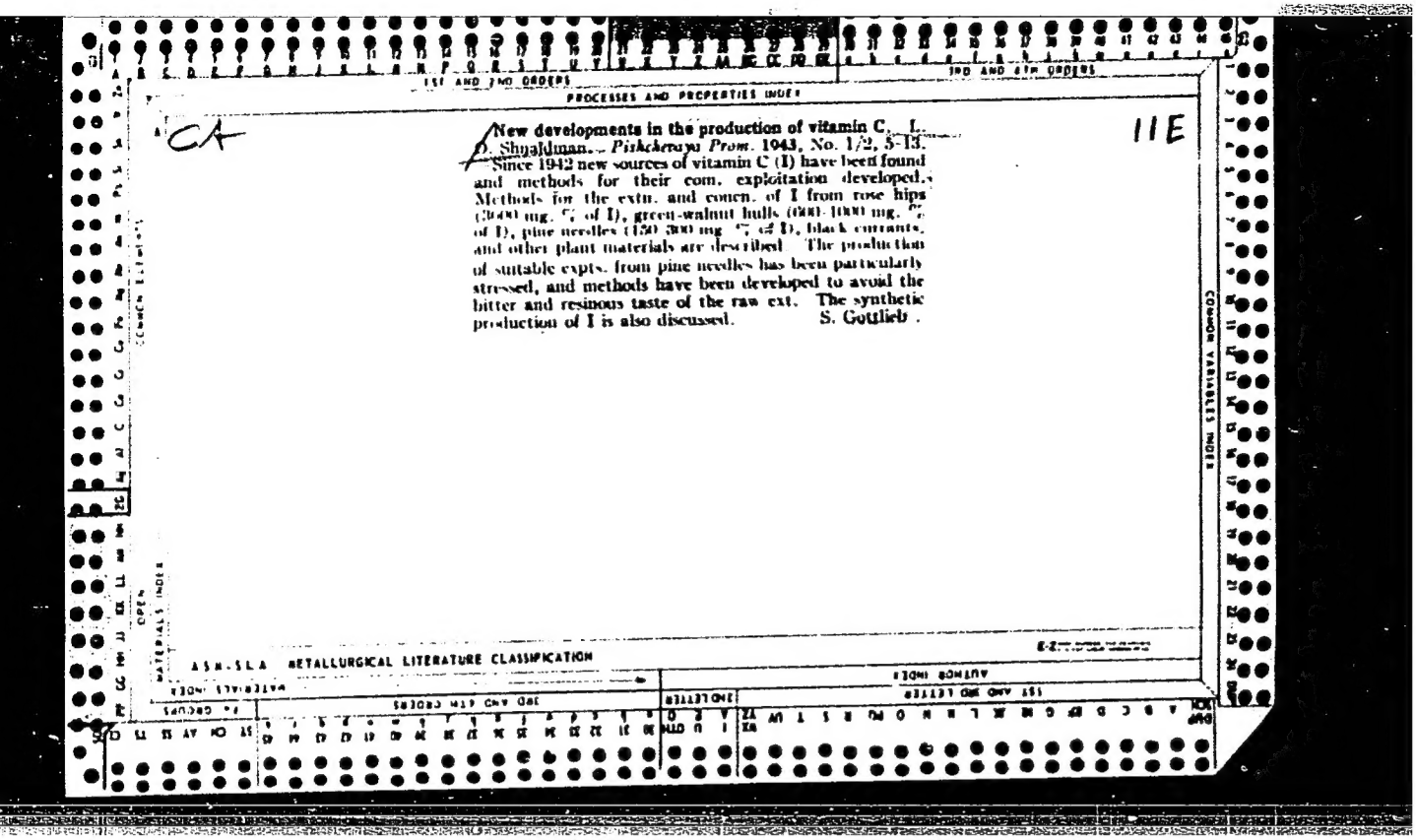
Edible beet sirups. N. E. Loginov and L. O. Shnaidman. *Trans. Central Sci. Research Inst. Sugar Ind.* (U. S. S. R.) No. 13, 57–68(1933).—Regular thick beet juice treated with 1 2% of Norit after inversion gives an edible sirup of high quality. A thick beet juice, carefully purified and sulfited after second carbonation, is treated with 1–2% of Norit, to decrease the color and remove beet flavor. Fifty % of the thick juice is inverted in a copper vacuum pan at 90–92° in 2–2.5 hrs. at p_H 3.5–4.0. One % of HCl on the wt. of sirup is used. The other 50% of sirup is evapd. in a vacuum pan up to 70°Brix at 80°. The inverted part of the sirup is neutralized with soda to p_H 5.5, and mixed in the vacuum pan with the evapd. sirup. Then all is boiled to 76–78°Brix at a low temp. The contact of the inverted sirup with iron app. before neutralizing with soda must be avoided. Rvapd. sirup must be quickly cooled. A no. of tables and descriptions of expts. are given.

V. R. Balkow



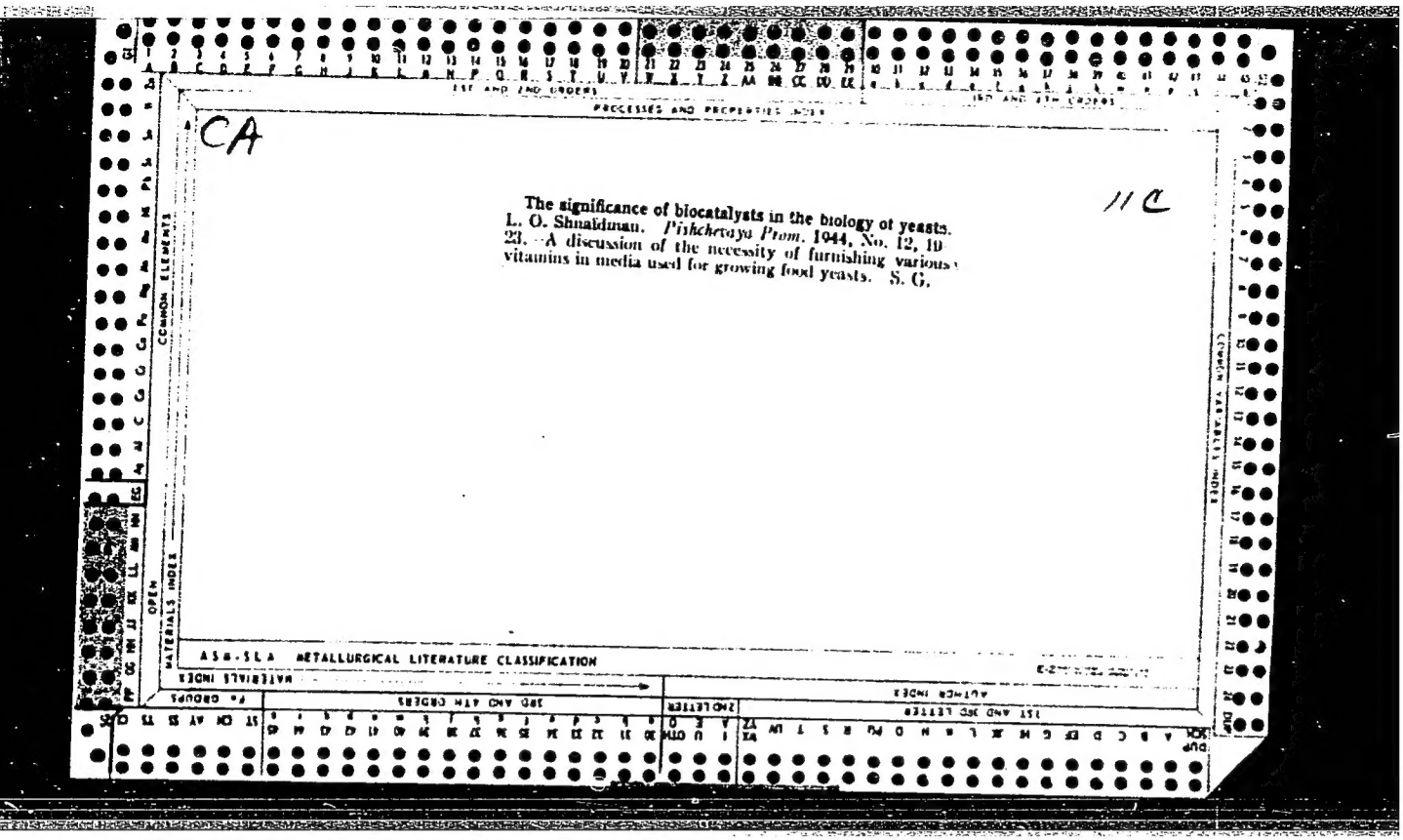
PROCESSES AND PROPERTIES INDEX

CA

11E

New developments in the production of vitamin C. L. D. Shvaldman. *Pishchevaya Prom.* 1943, No. 1/2, 5-13. Since 1942 new sources of vitamin C (I) have been found and methods for their com. exploitation developed. Methods for the extn. and concn. of I from rose hips (2000 mg. % of I), green-walnut hulls (600-1000 mg. % of I), pine needles (150-300 mg. % of I), black currants, and other plant materials are described. The production of suitable expts. from pine needles has been particularly stressed, and methods have been developed to avoid the bitter and resinous taste of the raw ext. The synthetic production of I is also discussed. S. Gottlieb.

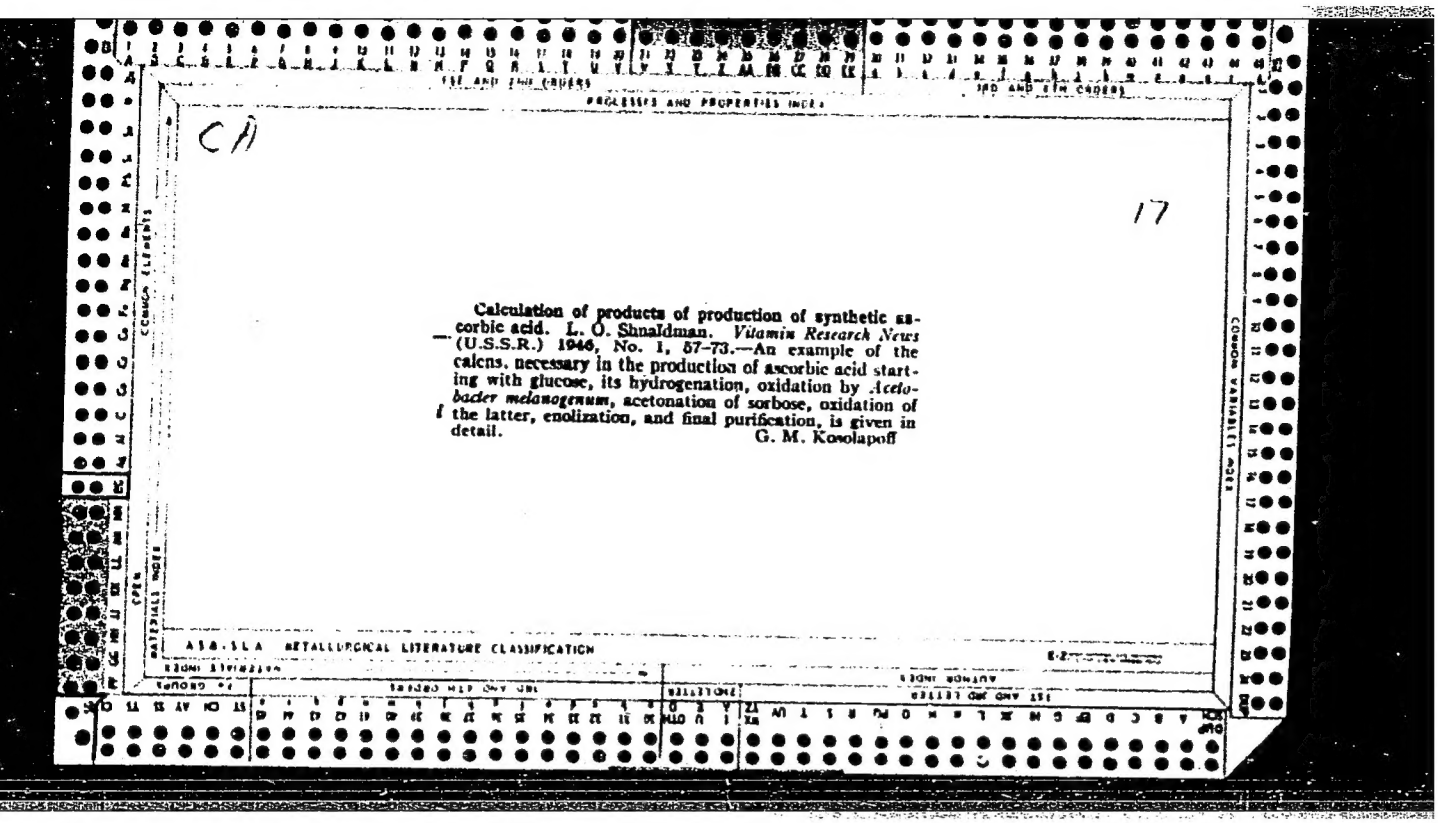
ASM-SLA METALLURGICAL LITERATURE CLASSIFICATION



CA

11C

The significance of biocatalysts in the biology of yeasts. L. O. Shnaiderman. *Pishchevaya Prom.* 1944, No. 12, 19-23.—A discussion of the necessity of furnishing various vitamins in media used for growing food yeasts. S. G.



CH

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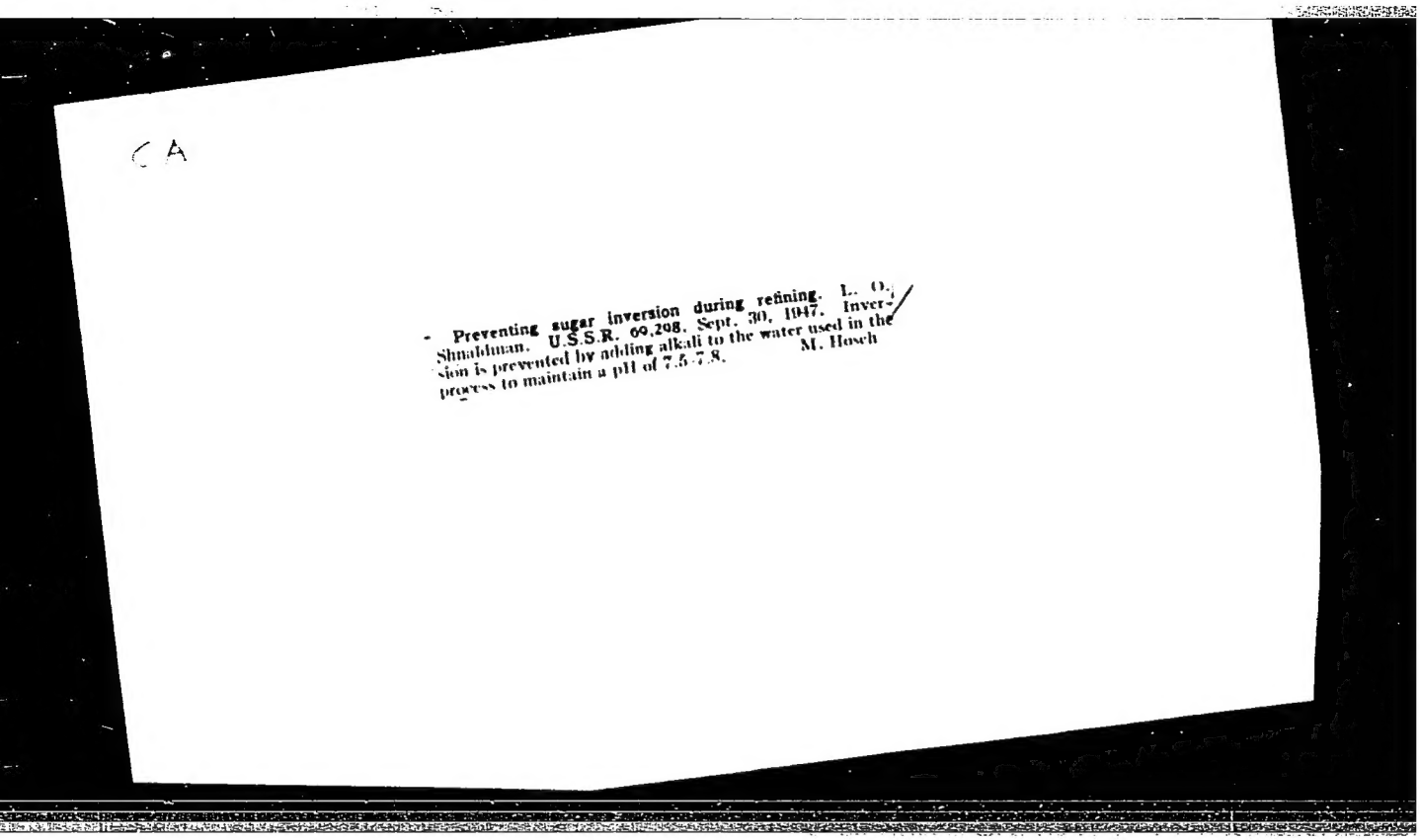
Calculation of products of production of synthetic ascorbic acid. L. O. Shnaldman. *Vitamin Research News* (U.S.S.R.) **1946**, No. 1, 57–73.—An example of the calcns. necessary in the production of ascorbic acid starting with glucose, its hydrogenation, oxidation by *Acetobacter melanogenum*, acetonation of sorbose, oxidation of the latter, enolization, and final purification, is given in detail.

G. M. Kosolapoff



CA

Preventing sugar inversion during refining. L. O. Shnaidman. U.S.S.R. 69,208. Sept. 30, 1947. Inversion is prevented by adding alkali to the water used in the process to maintain a pH of 7.5–7.8. M. Hosch